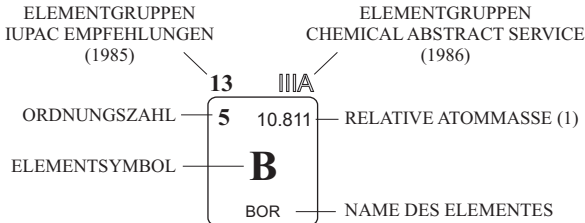


# PERIODENSYSTEM DER ELEMENTE

<http://www.periodni.com/de/>

PERIODEN	GRUPPE																18						
	1	2		3-10										11-12		13	14	15	16	17	VIIIA		
	IA	IIA		IIIB-VIIB										VIII		IB	IIB	IIIA	IVA	VA	VIA	VIIA	VIIIA
	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18					
1	1 1.0079 <b>H</b> WASSERSTOFF															5 10.811 <b>B</b> BOR	6 12.011 <b>C</b> KOHLENSTOFF	7 14.007 <b>N</b> STICKSTOFF	8 15.999 <b>O</b> SAUERSTOFF	9 18.998 <b>F</b> FLUOR	10 20.180 <b>Ne</b> HELIUM		
2	3 6.941 <b>Li</b> LITHIUM	4 9.0122 <b>Be</b> BERYLLIUM														13 26.982 <b>Al</b> ALUMINIUM	14 28.086 <b>Si</b> SILIZIUM	15 30.974 <b>P</b> PHOSPHOR	16 32.065 <b>S</b> SCHWEFEL	17 35.453 <b>Cl</b> CHLOR	18 39.948 <b>Ar</b> ARGON		
3	11 22.990 <b>Na</b> NATRIUM	12 24.305 <b>Mg</b> MAGNESIUM		3 44.956 <b>Sc</b> SCANDIUM	4 47.867 <b>Ti</b> TITAN	5 50.942 <b>V</b> VANADIUM	6 51.996 <b>Cr</b> CHROM	7 54.938 <b>Mn</b> MANGAN	8 55.845 <b>Fe</b> EISEN	9 58.933 <b>Co</b> KOBALT	10 58.693 <b>Ni</b> NICKEL	11 63.546 <b>Cu</b> KUPFER	12 65.409 <b>Zn</b> ZINK	31 69.723 <b>Ga</b> GALLIUM	32 72.64 <b>Ge</b> GERMANIUM	33 74.922 <b>As</b> ARSEN	34 78.96 <b>Se</b> SELEN	35 79.904 <b>Br</b> BROM	36 83.798 <b>Kr</b> KRYPTON				
4	19 39.098 <b>K</b> KALIUM	20 40.078 <b>Ca</b> CALCIUM	21 44.956 <b>Sc</b> SCANDIUM	22 47.867 <b>Ti</b> TITAN	23 50.942 <b>V</b> VANADIUM	24 51.996 <b>Cr</b> CHROM	25 54.938 <b>Mn</b> MANGAN	26 55.845 <b>Fe</b> EISEN	27 58.933 <b>Co</b> KOBALT	28 58.693 <b>Ni</b> NICKEL	29 63.546 <b>Cu</b> KUPFER	30 65.409 <b>Zn</b> ZINK	31 69.723 <b>Ga</b> GALLIUM	32 72.64 <b>Ge</b> GERMANIUM	33 74.922 <b>As</b> ARSEN	34 78.96 <b>Se</b> SELEN	35 79.904 <b>Br</b> BROM	36 83.798 <b>Kr</b> KRYPTON					
5	37 85.468 <b>Rb</b> RUBIDIUM	38 87.62 <b>Sr</b> STRONTIUM	39 88.906 <b>Y</b> YTTRIUM	40 91.224 <b>Zr</b> ZIRKON	41 92.906 <b>Nb</b> NIOB	42 95.94 <b>Mo</b> MOLYBDÄN	43 (98) <b>Tc</b> TECHNETIUM	44 101.07 <b>Ru</b> RUTHENIUM	45 102.91 <b>Rh</b> RHODIUM	46 106.42 <b>Pd</b> PALLADIUM	47 107.87 <b>Ag</b> SILBER	48 112.41 <b>Cd</b> KADMIUM	49 114.82 <b>In</b> INDIUM	50 118.71 <b>Sn</b> ZINN	51 121.76 <b>Sb</b> ANTIMON	52 127.60 <b>Te</b> TELLUR	53 126.90 <b>I</b> IOD	54 131.29 <b>Xe</b> XENON					
6	55 132.91 <b>Cs</b> CÄSIUM	56 137.33 <b>Ba</b> BARIUM	57-71 <b>La-Lu</b> Lanthaniden	72 178.49 <b>Hf</b> HAFNIUM	73 180.95 <b>Ta</b> TANTAL	74 183.84 <b>W</b> WOLFRAM	75 186.21 <b>Re</b> RHENIUM	76 190.23 <b>Os</b> OSMIUM	77 192.22 <b>Ir</b> IRIDIUM	78 195.08 <b>Pt</b> PLATIN	79 196.97 <b>Au</b> GOLD	80 200.59 <b>Hg</b> QUECKSILBER	81 204.38 <b>Tl</b> THALLIUM	82 207.2 <b>Pb</b> BLEI	83 208.98 <b>Bi</b> BISMUT	84 (209) <b>Po</b> POLONIUM	85 (210) <b>At</b> ASTAT	86 (222) <b>Rn</b> RADON					
7	87 (223) <b>Fr</b> FRANCIUM	88 (226) <b>Ra</b> RADIUM	89-103 <b>Ac-Lr</b> Actiniden	104 (267) <b>Rf</b> RUTHERFORDIUM	105 (268) <b>Db</b> DUBNIUM	106 (271) <b>Sg</b> SEABORGIUM	107 (272) <b>Bh</b> BOHRIUM	108 (277) <b>Hs</b> HASSIUM	109 (276) <b>Mt</b> MEITNERIUM	110 (281) <b>Ds</b> DARMSTADIUM	111 (280) <b>Rg</b> ROENTGENIUM												



## LANTHANIDEN

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57 138.91 <b>La</b> LANTHAN	58 140.12 <b>Ce</b> CER	59 140.91 <b>Pr</b> PRASEODYM	60 144.24 <b>Nd</b> NEODYM	61 (145) <b>Pm</b> PROMETHIUM	62 150.36 <b>Sm</b> SAMARIUM	63 151.96 <b>Eu</b> EUROPIUM	64 157.25 <b>Gd</b> GADOLINIUM	65 158.93 <b>Tb</b> TERBIUM	66 162.50 <b>Dy</b> DYSPROSIUM	67 164.93 <b>Ho</b> HOLMIUM	68 167.26 <b>Er</b> ERBIUM	69 168.93 <b>Tm</b> THULIUM	70 173.04 <b>Yb</b> YTTERBIUM	71 174.97 <b>Lu</b> LUTETIUM
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## ACTINIDEN

89 (227) <b>Ac</b> ACTINIUM	90 232.04 <b>Th</b> THORIUM	91 231.04 <b>Pa</b> PROTACTINIUM	92 238.03 <b>U</b> URAN	93 (237) <b>Np</b> NEPTUNIUM	94 (244) <b>Pu</b> PLUTONIUM	95 (243) <b>Am</b> AMERICIUM	96 (247) <b>Cm</b> CURIUM	97 (247) <b>Bk</b> BERKELIUM	98 (251) <b>Cf</b> CALIFORNIUM	99 (252) <b>Es</b> EINSTEINIUM	100 (257) <b>Fm</b> FERMIUM	101 (258) <b>Md</b> MENDELEVIUM	102 (259) <b>No</b> NOBELIUM	103 (262) <b>Lr</b> LAWRENCIUM
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(1) Pure Appl. Chem., 78, No. 11, 2051-2066 (2006)

Die relative atommasse wird auf fünf Stellen angezeigt. Für Elemente ohne stabile Isotope ist die Atommasse des stabilsten Isotops in Klammern gezeigtesotope. Drei dieser Elemente (Th, Pa und U) spielen eine bedeutende Rolle aufgrund ihrer Häufigkeit in der Erdkruste und ihre Atomgewichte und werden deshalb aufgelistet.