

# PERIODIC TABLE OF THE ELEMENTS

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PERIOD	GROUP NUMBERS IUPAC RECOMMENDATION (1985)																GROUP NUMBERS CHEMICAL ABSTRACT SERVICE (1986)																		
	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17
1	1 <b>H</b> HYDROGEN																	2 <b>He</b> HELIUM																	
2	3 <b>Li</b> LITHIUM	4 <b>Be</b> BERYLLIUM																	5 <b>B</b> BORON	6 <b>C</b> CARBON	7 <b>N</b> NITROGEN	8 <b>O</b> OXYGEN	9 <b>F</b> FLUORINE	10 <b>Ne</b> NEON											
3	11 <b>Na</b> SODIUM	12 <b>Mg</b> MAGNESIUM																	13 <b>Al</b> ALUMINIUM	14 <b>Si</b> SILICON	15 <b>P</b> PHOSPHORUS	16 <b>S</b> SULPHUR	17 <b>Cl</b> CHLORINE	18 <b>Ar</b> ARGON											
4	19 <b>K</b> POTASSIUM	20 <b>Ca</b> CALCIUM	21 <b>Sc</b> SCANDIUM	22 <b>Ti</b> TITANIUM	23 <b>V</b> VANADIUM	24 <b>Cr</b> CHROMIUM	25 <b>Mn</b> MANGANESE	26 <b>Fe</b> IRON	27 <b>Co</b> COBALT	28 <b>Ni</b> NICKEL	29 <b>Cu</b> COPPER	30 <b>Zn</b> ZINC	31 <b>Ga</b> GALLIUM	32 <b>Ge</b> GERMANIUM	33 <b>As</b> ARSENIC	34 <b>Se</b> SELENIUM	35 <b>Br</b> BROMINE	36 <b>Kr</b> KRYPTON																	
5	37 <b>Rb</b> RUBIDIUM	38 <b>Sr</b> STRONTIUM	39 <b>Y</b> YTTRIUM	40 <b>Zr</b> ZIRCONIUM	41 <b>Nb</b> NIOBIUM	42 <b>Mo</b> MOLYBDENUM	43 <b>Tc</b> TECHNETIUM	44 <b>Ru</b> RUTHENIUM	45 <b>Rh</b> RHODIUM	46 <b>Pd</b> PALLADIUM	47 <b>Ag</b> SILVER	48 <b>Cd</b> CADMIUM	49 <b>In</b> INDIUM	50 <b>Sn</b> TIN	51 <b>Sb</b> ANTIMONY	52 <b>Te</b> TELLURIUM	53 <b>I</b> IODINE	54 <b>Xe</b> XENON																	
6	55 <b>Cs</b> CAESIUM	56 <b>Ba</b> BARIUM	57-71 <b>La-Lu</b> Lanthanide	72 <b>Hf</b> HAFNIUM	73 <b>Ta</b> TANTALUM	74 <b>W</b> TUNGSTEN	75 <b>Re</b> RHENIUM	76 <b>Os</b> OSMIUM	77 <b>Ir</b> IRIDIUM	78 <b>Pt</b> PLATINUM	79 <b>Au</b> GOLD	80 <b>Hg</b> MERCURY	81 <b>Tl</b> THALLIUM	82 <b>Pb</b> LEAD	83 <b>Bi</b> BISMUTH	84 <b>Po</b> POLONIUM	85 <b>At</b> ASTATINE	86 <b>Rn</b> RADON																	
7	87 <b>Fr</b> FRANCIUM	88 <b>Ra</b> RADIUM	89-103 <b>Ac-Lr</b> Actinide	104 <b>Rf</b> RUTHERFORDIUM	105 <b>Db</b> DUBNIUM	106 <b>Sg</b> SEABORGIUM	107 <b>Bh</b> BOHRIUM	108 <b>Hs</b> HASSIUM	109 <b>Mt</b> MEITNERIUM	110 <b>Ds</b> DARMSTADIUM	111 <b>Rg</b> ROENTGENIUM																								

## LANTHANIDE

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57 <b>La</b> LANTHANUM	58 <b>Ce</b> CERIUM	59 <b>Pr</b> PRASEODYMIUM	60 <b>Nd</b> NEODYMIUM	61 <b>Pm</b> PROMETHIUM	62 <b>Sm</b> SAMARIUM	63 <b>Eu</b> EUROPIUM	64 <b>Gd</b> GADOLINIUM	65 <b>Tb</b> TERBIUM	66 <b>Dy</b> DYSPROSIUM	67 <b>Ho</b> HOLMIUM	68 <b>Er</b> ERBIUM	69 <b>Tm</b> THULIUM	70 <b>Yb</b> YTTTERBIUM	71 <b>Lu</b> LUTETIUM
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## ACTINIDE

89 <b>Ac</b> ACTINIUM	90 <b>Th</b> THORIUM	91 <b>Pa</b> PROTACTINIUM	92 <b>U</b> URANIUM	93 <b>Np</b> NEPTUNIUM	94 <b>Pu</b> PLUTONIUM	95 <b>Am</b> AMERICIUM	96 <b>Cm</b> CURIUM	97 <b>Bk</b> BERKELIUM	98 <b>Cf</b> CALIFORNIUM	99 <b>Es</b> EINSTEINIUM	100 <b>Fm</b> FERMIUM	101 <b>Md</b> MENDELEVIUM	102 <b>No</b> NOBELIUM	103 <b>Lr</b> LAWRENCIUM
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(1) Pure Appl. Chem., 78, No. 11, 2051-2066 (2006)

Relative atomic mass is shown with five significant figures. For elements have no stable nuclides, the value enclosed in brackets indicates the mass number of the longest-lived isotope of the element. However three such elements (Th, Pa, and U) do have a characteristic terrestrial isotopic composition, and for these an atomic weight is tabulated.